

### **chapter 10 chemical quantities pdf**

1 CHEMICAL QUANTITIES Chapter 10. 2 What is a mole? A unit of measurement in chemistry 1 mole of a substance =  $6.02 \times 10^{23}$  (Avagadro s number) representative particles of a substance Representative particle atoms, molecules (nonmetals), formula units (metal-nonmetal), ions.

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CHAPTER 10: Chemical Quantities BASICS:  $\hat{\text{€}}$  The basic unit that is used to determine the amount of a chemical substance is called a mole  $\hat{\text{€}}$  A mole(mol) of a substance is equivalent to  $6.02 \times 10^{23}$  particles of that substance  $\hat{\text{€}}$  The mole was founded by a scientist named Avagadro, and he decided to use the

### **CHAPTER 10: Chemical Quantities**

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### **chapter 10- chemical quantities | Mole (Unit) | Chemical**

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Chapter 10: Chemical Quantities Section 10.1: The mole, a measurement of matter. How do you measure the amount of matter One way is to simply count the number of ... A mole (mol) of a substance is  $6.02 \times 10^{23}$  representative particles of that substance and is the SI unit for measuring the amount of a substance.

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Chemical Quantities THE MOLE AND QUANTIFYING MATTER 10.1 The Mole: A ... Use the chemical formula to find the number of atoms in one molecule and multiply ... The quantities 1 mol and 22.4 L can be used in conversion factors that change moles to

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Chapter 10 Chemical Quantities. 259. 33. What is the empirical formula of a compound that is 27.3% C and 72.7% O? 34. A compound consisting of 56.38% phosphorus and 43.62% oxygen has a molar mass of 219.9 g/mol. Determine its molecular formula. D. Essay. Write a short essay for the following. 35.

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Balanced chemical equations tell us in what proportions on a mole basis substances combine; since the molar masses of C (s) and O<sub>2</sub> (g) are different, 1 g of O<sub>2</sub> could not represent the same number of moles as 1 g of C.

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Chapter 10 Chemical Quantities 91 SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287–296) This section defines the mole and explains how the mole is used to measure matter. It also teaches you how to calculate the mass of a mole of any substance. Measuring Matter (pages 287–289) 1.

### SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287–296)

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Chapter 9 . Chemical Quantities: Stoichiometry . Test Review . Define the following: 1. Stoichiometry . 2. Limiting Reactant . 3. Theoretical Yield . 4. Excess Reactant . 5. Actual Yield . 6. Percent Yield . 7. What do the coefficients in a balanced equation tell you? 8. When NaCl is prepared from 37 g of Na, and an excess of Cl

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Guided Notes Chapter 10: Chemical Quantities. Guided Notes Chapter 10: Chemical Quantities Name \_\_\_\_\_ 10.1 The Mole: A Measurement of Matter 1. What is the mole 2. What is Avogadro's number 3. What kinds of things can you count using the mole 4.

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Percent Composition and Chemical Formulas 11) A compound analyzed in a chemistry laboratory consists of 5.34 g of carbon, 0.42 g of hydrogen, and 47.08 g of chlorine.

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